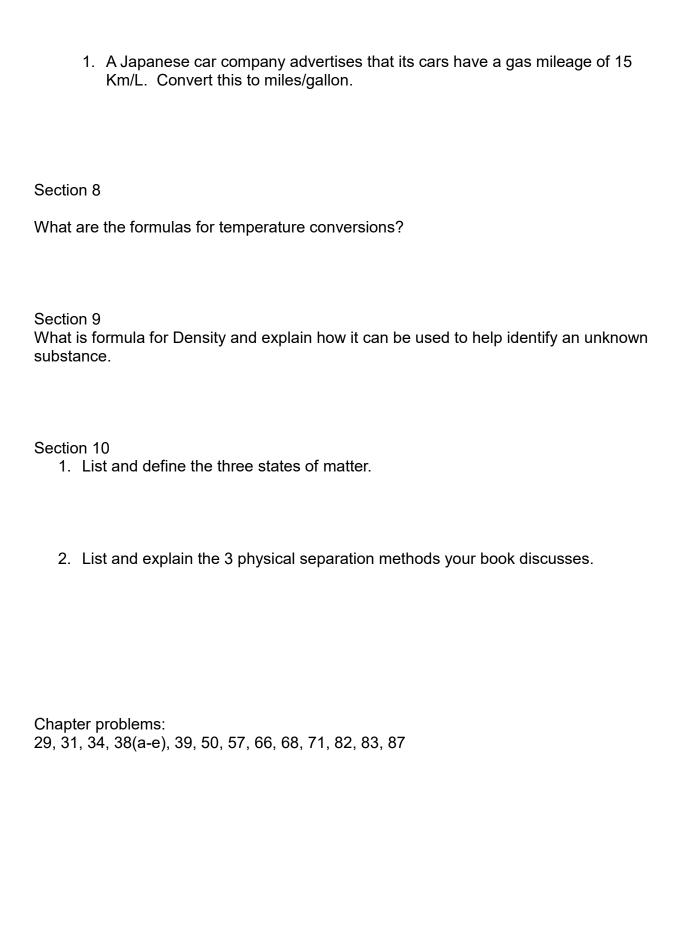
AP Chemistry Chapter 1 Reading Questions

Section 1

	What is one surprising Chem fact from the introduction?
	2. What can be used to view individual atoms?
	3. Compare macroscopic vs microscopic.
	4. Review question: List the 7 diatomic molecules.
	5. What two fundamental concepts of Chemistry are illustrated in section 1.1?
Section	on 2
1.	Compare qualitative and quantitative observations.
2.	What is another word for scientific theory?
3.	Compare scientific theory and law.
	Section 3

A quantitative measure always includes what 2 things?
2. 1 cm3 = ? mL
3. Compare mass and weight.
Section 4
1. How many numbers in a measurement are uncertain?
2. What is the rule for recording measurements to correct significant figures?
3. Compare accuracy and precision.
4. Compare random and systemic error.
Section 5 1. Make a set of rules for counting and using sig figs in calculations.
Section 6 X Section 7



Chapter 2 Reading questions

Section 2.2 1. List and describe the 3 fundamental chemical laws.
Section 2.3 1. List the 4 proponents of Dalton's theory.
2. What is Avogadro's hypothesis?
Section 2.4 Describe the contributions made by the following people to atomic theory: JJ Thomson
Becquerel
Rutherford

- 2.5
- 1. Define isotope
 - 2. What is "Z" and what does it mean. What is "A" and what does it mean.
 - 3. What is a covalent bond? And, the resulting collection of atoms is called?

4. What are the 4 ways of representing molecules?
5. How do ions form?
6. What is ionic bonding?
7. Define ionic solid.
8. Define polyatomic ion.
Section 2.7 1. List the properties of metals:
2.List the properties of nonmetals:
3. What are the 4 families with names
Section 2.8 1. What are the rules for naming type 1 binary ionic compounds:

- 2. Rules for naming type II Binary ionic compounds
- 3. Rules for naming oxyanions
- 4. Rules for naming binary covalent molecules
- 5. Rules for naming acids

Chapter 2 problems:

49, 51, 53, 56, 62, 64, 68, 70, 72, 73, 74, 75, 76, 80, 84, 86, 88

Chapter 3 Reading Questions

1.	Why do we count atoms by weighing?
2.	Describe how a mass spectrometer works.
3.	Define amu.
4.	Define Avogadro's #
5.	Define molar mass.
6.	Compare empirical and molecular formulas.
7.	Define stoichiometric quantities.
8.	Define Limiting reactant.
9.	Define Theoretical Yield

Chapter Problems: 38, 46, 48, 50, 51, 53, 55, 57, 59, 61, 68, 70, 76, 79, 82, 83, 85, 87, 96, 97, 102, 105, 108, 117, 124, 125